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CHAPTER 8 SOLUTIONS MANUAL

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Solutions Manual Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240–247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH_3 $\text{H}_2\text{C}=\text{CH}_2$ $\text{H}_2\text{C}=\text{C}=\text{CH}_2$ respectively, for single, double, and triple bonds. 2. H_2S H_2O H_2O_2 3.

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Section 8.4 – Polar Bonds and Molecules. Covalent bonds involve sharing electrons between atoms.

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When the atoms in the bond pull equally, the bonding electrons are shared equally, and the bond is nonpolar. When the atoms in the bond pull unequally, the bonding electrons are pulled closer to one atom, and the bond is polar.

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hydrogen's 1s orbital to form a covalent bond, called σ bond. The angle between each lobe assumes 109.47° since the sp^3 carbon is tetrahedral. sp^2 Hybrid Orbitals : If a carbon atom is in a double-bond environment, such as C_2H_4 (ethylene), the carbon is hybridized to

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Chapter 8. Covalent Compounds: Bonding Theories and ...

Lattice energies of alkali metal iodides.
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ionic radii sums for LiF and MgO are
2.01 and 2.06 Å, respectively, yet their

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lattice energies are 1030 and 3795
kJ/mol.

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In covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of noble gases.

- For example, a single hydrogen atom has one electron. But a pair of

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hydrogen atoms share electrons to form a covalent bond in a diatomic hydrogen molecule.

Chapter 8

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