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element in a compound is replaced by another element. One reactant is a compound and the other is oxygen.  
no reaction reaction SECTION 11.2 TYPES OF CHEMICAL REACTIONS (pages 330–339) Pearson Prentice Hall Chemistry Workbook

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involve changes in chemical. Energy in Reactions (page 50) 4. What is released or absorbed whenever chemical bonds form or are broken? 5. What do chemical reactions that absorb energy need ...

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Solution Formation and Chemical Reactions

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Decomposition chemical reaction is the reaction where only one compound

decomposes and results in two or more than two products.  $\text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbO} + \text{NO}_2 + \text{O}_2$

. In this equation, lead nitrate is being decomposed, which breaks down to form nitrogen dioxide, oxygen, and lead oxide.

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Reaction 1 : 50 ml of 2M HCl = 50 ml of 2M NaOH Moles of HCl = Moles of NaOH =  $0.05 \text{ L} \times 2 \text{ mol/L} = 0.1 \text{ moles}$  T = 12.86 C m (water ) =  $100 \text{ ml} \times 1.03 \text{ g/ml} = 103 \text{ g}$

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$q(\text{reaction}) = -m \cdot T \cdot C(\text{water}) = -103$  view the full answer

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